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Name _____
Period : 1 2 3 4 5 6

Elements and their Properties
19-1, 19-2, & 19-3 Section Review

Choose the term from the word list that best completes each statement. Write the term in the blank at the left of each statement.

- | | | |
|------------------------------------|-----------------------------------|-----------------------------------|
| diatomic
ductile
noble gases | halogens
hydrogen
nonmetals | malleable
metals
transition |
|------------------------------------|-----------------------------------|-----------------------------------|
- Hydrogen 1. _____ is the only nonmetal on the left side of the periodic table.
- transition 2. Metals in Groups 3 through 12 of the periodic table are called _____ elements.
- diatomic 3. Two atoms of the same element, when bonded, form a(n) _____ molecule.
- malleable 4. Aluminum can be made into foil and cans because it is _____.
- halogens 5. Fluorine, chlorine, bromine, and iodine are _____.
- noble gases 6. The elements in Group 18 are commonly known as the _____.
- nonmetals 7. The elements found on the right side of the periodic table are classified as _____.

Choose the answer that best completes each of the following sentences. Write the letter for that answer on the line to the left of each question.

- _____ 8. At room temperature, most metals are _____.
a. gases c. radioactive
b. liquids d. **solids**
- _____ 9. Hydrogen is grouped with the alkali metals because it _____.
a. does not readily form compounds
b. **has one electron in its outer energy level**
c. is a gas
d. is a metal

Write names of the following compounds

Pb ₃ (PO ₄) ₂	<u>lead(II) phosphate</u>	CoSO ₄	<u>Cobalt(II) sulfate</u>
Cu ₂ SO ₄	<u>copper(I) sulfate</u>	CuSO ₃	<u>Copper(II) sulfite</u>
Pb(ClO ₃) ₂	<u>lead(II) chlorate</u>	Hg(IO ₃) ₂	_____
Fe(ClO ₄) ₃	<u>Iron(III) perchlorate</u>	Fe ₂ (SO ₄) ₃	<u>Iron(III) sulfate</u>
Cr(OH) ₂	<u>Chromium(II) hydroxide</u>	KBrO ₃	<u>Potassium Bromate</u>
Cd(IO ₃) ₂	<u>Cadmium(II) Iodate</u>	FeSO ₄	<u>Iron(II) sulfate</u>
CaSO ₃	<u>Calcium sulfite</u>	BaCO ₃	<u>Barium carbonate</u>
KNO ₃	<u>Potassium Nitrate</u>	NH ₄ C ₂ H ₃ O ₂	<u>Ammonium Acetate</u>
Pb ₃ N ₂	<u>Lead(II) Nitride</u>	Ba ₃ P ₂	<u>Barium phosphide</u>
NaClO ₄	<u>Sodium perchlorate</u>	NaNO ₃	<u>Sodium Nitrate</u>
Mg(NO ₃) ₂	<u>magnesium Nitrate</u>	SbPO ₄	<u>Antimony(III) phosphate</u>
Li ₂ SO ₄	<u>lithium sulfate</u>	(NH ₄) ₂ CO ₃	<u>Ammonium carbonate</u>
Cu ₂ CrO ₄	<u>Copper(II) chromate</u>	NaHCO ₃	<u>Sodium Hydrogen carbonate</u>
Ca(ClO ₂) ₂	<u>Calcium chlorite</u>	Ba(NO ₂) ₂	<u>Barium Nitrite</u>
Ag ₂ CrO ₄	<u>Silver chromate</u>	Sn ₃ (PO ₄) ₂	<u>Tin(II) phosphate</u>
H ₂ SO ₄	<u>Hydrogen sulfate</u>	CaH ₂	<u>Calcium hydride</u>
AuClO ₄	<u>gold(I) Perchlorate</u>	Mg(OH) ₂	<u>Magnesium hydroxide</u>
KHSO ₄	<u>Potassium hydrogen sulfate</u>	SnCl ₂	<u>Stannous Chloride</u>
MnS	<u>manganese sulfide</u>	AgNO ₃	<u>Silver Nitrate</u>
Ni(NO ₃) ₂	<u>Nickel(II) Nitrate</u>	FeHPO ₄	Iron
PCl ₅	<u>phosphorus pentachloride</u>	N ₂ O ₂	<u>Nitrogen dioxide</u>
N ₂ O ₃	<u>dinitrogen trioxide</u>	N ₂ O	<u>dinitrogen monoxide</u>
CCl ₄	<u>Carbon tetrachloride</u>	N ₂ O ₅	<u>dinitrogen pentoxide</u>

Answers to worksheets

1.1 Carl's new experiments

1 Qualitative. They were qualitative because all observations were recorded using only words.

2 Carl thought that he should also do a quantitative analysis, that is, record the observations using numbers to measure how much the grass actually grows under different situations.

3 a Controls.
b Every experiment needs a control against which results can be compared.

4 Types of grass, amount of grass.

5 The amount of water was kept the same.

6 The amount of sunlight.

7 The amount of water.

8 The soil would have had some water in it on Day 1 and would take a few days to dry out.

9 Measurements to be taken at the same time of day. Same ruler to be used and measured from the same place. Others as appropriate.

10 Repeating measurements is a good way to improve accuracy. Once a collection of different measurements is taken, an average can be obtained.

11 a

Carl's experiments

Days	1A	1B	2C	2D
0	12	12	12	12
1	13	13	11	10
2	14	13	10	9
3	15	14	9	8
4	16	14	9	7
5	16	14	8	6

b 1A 16 or 17 mm
1B 14 or 15 mm
2C 8 or 9 mm
2D 6 or 7 mm

Source: Focus 1 Worksheet, Basic Anatomy of Science (Student Resource 1), Biology of Science Institute Group, Inc. 2005.
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